Our para-azoxyanisol sample was purchased from Eastman Kodak company; we did not make any purification before or during its handling, however the transition temperatures were checked after each measurement: in particular, their values were not changed after experiments up to 3 kbar; this indicates that there is no substantial degradation in the corresponding temperature range.

3. Phase Diagram

The effect of pressure on the transition temperatures is shown in Fig. 1; clearly, this is a very large effect: common slopes for the

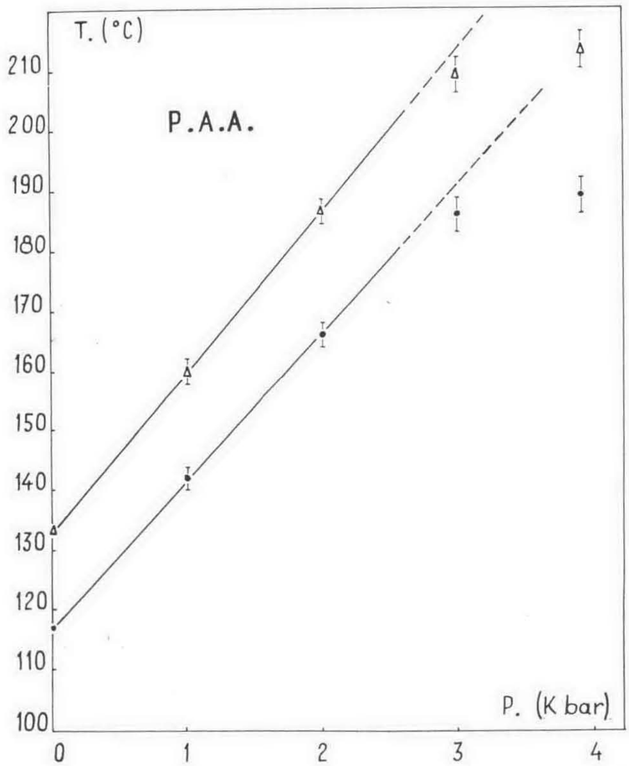

Figure 1. Variation with pressure of the temperatures of the solid–nematic (dots) and nematic–isotropic (triangles) transitions in T, P. The slopes are roughly coherent with other existing thermodynamic data, if one applies the Clausius–Clapeyron equation (see Table 2). The non linearity above 3 kb is due to thermal degradation of the sample: the points at 4 kb are not significant

TABLE 1

	Slope of solid–nematic transition line: $\frac{dT_M}{dP}$	Slope of nematic–isotropic transition line: $\frac{dT_c}{dP}$
G. A. Hulett[7]	32°/kbar up to 300 bars	48.6°/kbar up to 300 bars
G. Puschin and W. Grebenschtschikov[8]	25.6°/kbar up to 2 kbar	39.4°/kbar up to 1 kbar
J. Robberecht[9]	32°/kbar up to 850 bars	48°/kbar up to 935 bars
This work	a) 24.5°/kbar up to 3 kbar b) 23.7°/kbar up to 3 kbar[10]	27°/kbar up to 3 kbar 28.7°/kbar up to 3 kbar[10]

TABLE 2

	Relative jump in specific volume at melting point $\left(\frac{\Delta V}{V}\right)_M$	Relative jump in specific volume at clearing point $\left(\frac{\Delta V}{V}\right)_c$
W. Maier and A. Saupe[18]		0.30% 0.35% " extrapoled " value
E. McLaughlin, A. Shakespeare and R. Ubbelohde[11]	11.03%	0.36%
	Latent heat at melting point: $(\Delta H)_M$	Latent heat at clearing point: $(\Delta H)_c$
R. Schenck[12]		735 joules/mole
C. Kreutzer and W. Kast[13]		1730 joules/mole
H. Arnold[14]		600 joules/mole
I. G. Chistyakov[15]	29.570 joules/mole	574 joules/mole
E. M. Barral, R. S. Porter and J. F. Johnson[16]	30.200 joules/mole	735 joules/mole
L. C. Chow and D. E. Martire[29]		760 joules/mole
This work, from Clausius–Clapeyron equation	a) 35.000 joules/mole b) 36.000 joules/mole[10]	1.080 joules/mole 1.000 joules/mole[10]